

Name _____

Periodic Table Lab Questions:
Data Tables

1.) Atomic Mass for Period 3

Element	Atomic Mass

2.) Atomic Number for Period 2

Element	Atomic Number

3.) Atomic Number for group 2

Element	Atomic #
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4.) Atomic size (radius) for Period 2

Element	Atomic Radius

5.) Atomic size (radius) for group 1

Element	Atomic Size

6.) Ionization Energy for group 17

Element	Ionization Energy
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7.) Ionization Energy for Period 3

Element	Ionization Energy

8.) Electronegativity for group 13

Element	Electronegativity
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9.) Electronegativity for period 4

Element	Electronegativity

1.) Define each of the following terms:

Electronegativity: _____

Ionization Energy: _____

Atomic Radius: _____

Atomic Mass: _____

Atomic Number: _____

2.) On the periodic table, what represents a family, vertical columns or horizontal rows?

3.) On the periodic table, what represents a group, vertical columns or horizontal rows?

4.) Looking at the graphs that you made describe what happens in each of the following scenarios with the word increase, decrease, or remains the same.

What happens to the atomic number as you go down a group? _____

What happens to the atomic radius as you go across a period? _____

What happens to the atomic radius as you go down a group? _____

What happens to ionization energy as you go down a group? _____

What happens to ionization energy as you go across a period? _____

What happens to electronegativity as you go down a group? _____

What happens to electronegativity as you go across a period? _____

5.) The best relationship is the one with the straightest line. Which is the best way for organizing the elements in the periodic table, atomic mass or atomic number?









